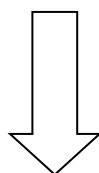


Review Sheet Chapter 19 Acids/Bases 122 Mrs Sanford		Name:
Properties of Acids-		
Properties of Bases-		
Monoprotic-		Diprotic-
Arrhenius Acid-		
Arrhenius Base-		
Lewis diagram and structural diagram of hydronium ion-		Amphoteric-
Bronsted Lowry Acid:		Bronsted Lowry Base:
Bronsted/Lowry -give a reaction example with A and B labeled-		
Conjugate acid /base pairs-Give an example:		
The stronger the acid,the_____its conjugate base.(when looking at the acid/base table)		
Example of a Lewis acid and a Lewis base:		
Self-Ionization of water:		
What does Kw stand for and give an equation for it with the value of the constant:		
In a basic solution:		In an acidic solution:
pH=	pOH=	14=
Strong Acid:		Weak Acid:
How does the value of Ka get affected by strength of the acid?		
Ka=	Kb=	
Calculating a Dissociation constant using an ICE table: There is a calculation of this type below. P 610 good review		
Know how to use the indicator table.		
Five steps for predicting acid/base reactions 1-		
2-		
3-		
4-		
5-		
Draw a sketch and label a pH curve with a base as a titrant and two quantitative reactions.		Draw a sketch and label a pH curve with An acid as a titrant and one quantitative Reaction.



(Label Ph endpoint, equivalence point, point of neutralization on your sketch graphs)

What does a strong dilute acid mean?	What does a strong concentrated acid mean?	
Acid base stoichiometry How do you find volume of titrant?	Why is a balanced equation important?	Write an example of a neutralization reaction involving coefficients other than 1.
Do. P 625 #45 to 50 55 to 57 58 to 61 63, 65 p 626 74 a,c, 75a,c,85, p 610 22, 23(ICE calculations), Give an example for a 5 step with \rightarrow , $>50\%$, and $<50\%$ (an example for each)		